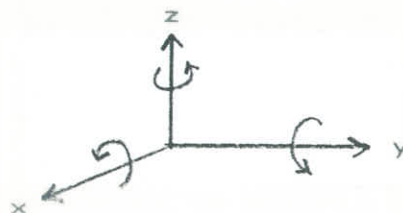


SYMMETRY ANALYSIS FOR H₂O

Water belongs to the C_{2v} symmetry group and has the following symmetry elements: E, C₂^z, σ_{xz}, and σ_{yz}. Its character table is shown below.

C _{2v}	E	C ₂ ^z	σ _{xz}	σ _{yz}	h = 4
A ₁	1	1	1	1	z, x ² , y ² , z ²
A ₂	1	1	-1	-1	xy, R _z
B ₁	1	-1	1	-1	x, xz, R _y
B ₂	1	-1	-1	1	y, yz, R _x
Γ _{tot}	9	-1	1	3	

The C_{2v} symmetry group has four symmetry elements and four associated symmetry operations. It can be thought of as a four dimensional space with the A₁, A₂, B₁, and B₂ irreducible representations playing the role of unit vectors of vector algebra. The irreducible representations span the space and any other vector or representation in that space can be written as a linear combination of them. For simple groups like this the character table can be generated by examining how translations along the x-, y-, and z-axes and rotations about these axes transform under the symmetry operations of the group. Using the figure shown below you should be able to confirm the designations shown in the right-hand column of the character table. That is that translation in the z-direction transforms like A₁, rotation about the z-axis transforms like A₂, translation in the x-direction and rotation about the y-axis have symmetry properties represented by B₁, and translation in the y-direction and rotation about the x-axis have B₂ symmetry. In doing this note (see Atkins) that you record +1 if something is transformed into itself, -1 if it is transformed into minus itself, and 0 if it is transformed into something else under the symmetry operations of the group.



Note that just like the unit vectors in the cartesian coordinate system, these irreducible representations are orthogonal. That is they have zero overlap.

$$A_1 \cdot B_2 = [(1)(1) + (1)(-1) + (1)(-1) + (1)(1)]/4 = 0$$

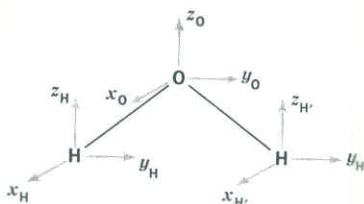
$$A_2 \cdot B_1 = [(1)(1) + (1)(-1) + (-1)(1) + (-1)(-1)]/4 = 0$$

Dividing by 4, the order of the group, normalizes the calculation. We can also show that all of the irreducible representations are normalized or have an overlap of 1.

$$A_2 \cdot A_2 = [(1)(1) + (1)(1) + (-1)(-1) + (-1)(-1)]/4 = 1$$

Note that these vector operations are exactly equivalent to evaluating quantum mechanical overlap integrals $\int \psi_1 \psi_2 d\tau$.

Since H₂O is a triatomic molecule it has a total of 9 degrees of freedom - three for each atom in the molecule. These are the x-, y-, and z-coordinates of the individual atoms and are shown as small vectors located on each of the atoms in the figure below.



A symmetry analysis for water begins by determining how these 9 coordinates behave under the symmetry operations of the C_{2v} group. You should be able to show that this generates the reducible representation, Γ_{tot} , which is given in the last row of the character table shown above. Γ_{tot} can also be calculated as $\Gamma_{\text{uma}} \cdot (\Gamma_x + \Gamma_y + \Gamma_z)$, where Γ_{uma} is the behavior of the atoms under the symmetry operations of the group. Of the 9 degrees of freedom possessed by the water molecule, three are for translation of the center of mass in the x-, y-, and z-directions, and three are related to rotation about the x-, y-, and z-axes. This leaves three vibrational degrees of freedom. To determine the symmetry of the vibrational modes we decompose Γ_{tot} into the unit vectors or irreducible representations of the C_{2v} character table. This involves taking the dot product of Γ_{tot} with each of the irreducible representations of the C_{2v} symmetry.

$$\Gamma_{\text{tot}} \cdot A_1 = [(9)(1) + (-1)(1) + (1)(1) + (3)(1)]/4 = 3$$

$$\Gamma_{\text{tot}} \cdot A_2 = [(9)(1) + (-1)(1) + (1)(-1) + (3)(-1)]/4 = 1$$

$$\Gamma_{\text{tot}} \cdot B_1 = [(9)(1) + (-1)(-1) + (1)(1) + (3)(-1)]/4 = 2$$

$$\Gamma_{\text{tot}} \cdot B_2 = [(9)(1) + (-1)(-1) + (1)(-1) + (3)(1)]/4 = 3$$

This procedure has revealed that the reducible representation, Γ_{tot} , is composed of the following irreducible representations:

$$\Gamma_{\text{tot}} = 3A_1 + A_2 + 2B_1 + 3B_2$$

From the right side of the character table we see that translation and rotation have the following symmetry properties:

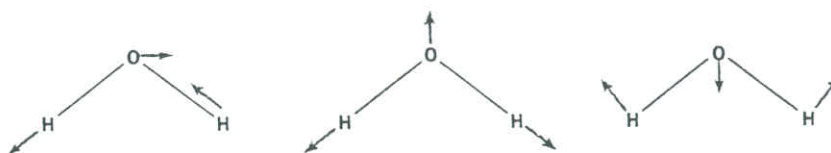
$$\Gamma_{\text{trans}} = A_1 + B_1 + B_2$$

$$\Gamma_{\text{rot}} = A_2 + B_1 + B_2$$

From this information we can determine the symmetry properties of the vibrational modes of the water molecule.

$$\Gamma_{\text{vib}} = \Gamma_{\text{tot}} - \Gamma_{\text{trans}} - \Gamma_{\text{rot}} = 2A_1 + B_2$$

There are three vibrational fundamentals. And, since there are two bonds there will be two stretching vibrations and one bending vibration. To determine which symmetry classification the bend belongs to examine how the H-O-H bond angle transforms under the symmetry operations of the C_{2v} group.



The vibrational modes are shown above. Convince yourself that their symmetry properties are captured in the table shown below.

C_{2v}	E	C_2^z	σ_{xz}	σ_{yz}	
Γ_{bend}	1	1	1	1	A_1
Γ_{stretch}	1	1	1	1	A_1
Γ_{stretch}	1	-1	-1	1	B_2

Whether any of these modes is active in the infrared region of the spectrum is another question. This is determined by evaluating the quantum mechanical transition probability integral,

$$P = e \int \Psi^f(\text{vib}) \hat{q} \Psi^i(\text{vib}) d\tau$$

where q is the spatial coordinate and is either x , y , or z . If this integral is non-zero the transition is allowed. Group theory enables us to determine if the integral is non-zero as follows. We evaluate the direct product (note the similarity to the quantum mechanical transition probability integral) $\Gamma_{\text{vib}}^f \cdot \Gamma_q \cdot \Gamma_{\text{vib}}^i$ and if it turns out to be equal to, or contain, the irreducible representation A_1 then the transition is allowed. If the direct product isn't equal to or contain A_1 the transition is forbidden.

Γ_{vib}^i , the representation for the ground state is always equal to A_1 and Γ_{vib}^f has the

symmetry of the vibrational mode being excited, which in our case is either A_1 or B_2 . Thus we can see that a vibrational mode will be infrared active if it belongs to the same symmetry species as one of the cartesian coordinates. You should verify that this is correct and also be able to show that all three vibrational modes of the water molecule are infrared active.

Group theory is also very helpful in constructing molecular orbitals from linear combinations of atomic orbitals. For the water molecule the atomic orbitals are the hydrogen 1s orbitals and the 2s and the three 2p orbitals on the oxygen atom. We proceed by examining how the orbitals transform under the symmetry operations of the group. This is shown in the table below.

C_{2v}	E	C_2^z	σ_{xz}	σ_{yz}
H_1	H_1	H_2	H_2	H_1
O_{2s}	O_{2s}	O_{2s}	O_{2s}	O_{2s}
O_{2px}	O_{2px}	$-O_{2px}$	O_{2px}	$-O_{2px}$
O_{2py}	O_{2py}	$-O_{2py}$	$-O_{2py}$	O_{2py}
O_{2pz}	O_{2pz}	O_{2pz}	O_{2pz}	O_{2pz}
A_1	1	1	1	1
A_2	1	1	-1	-1
B_1	1	-1	1	-1
B_2	1	-1	-1	1

To find the molecular orbital with A_1 symmetry take the sum of the dot product of the A_1 irreducible representation with each of the representations in the top half of the table. This is $2(H_1 + H_2) + 4O_{2s} + 4O_{2pz}$. This indicates that the molecular orbital will consist of a linear combination of four of the six orbitals shown in the table above.

$$\Psi(A_1) = c_h[\psi(H_1) + \psi(H_2)] + c_o\psi(O_{2s}) + c_o'\psi(O_{2pz})$$

When this is done for A_2 symmetry it is found that the sum is zero, indicating that there is no molecular orbital with A_2 symmetry. Repeating this process for B_1 symmetry yields $4O_{2px}$. This means $\Psi(B_1) = \Psi(O_{2px})$ and we say that the oxygen $2p_x$ orbital is non-bonding because there are no hydrogen orbitals with the same symmetry. This is easily seen by inspection of the water molecule when placed in a cartesian coordinate system as shown in the figure above.

For B_2 symmetry the result is $2(H_1 - H_2) + 4O_{2py}$. This is clearly another bonding molecular orbital.

$$\Psi(B_2) = c_h'[\psi(H_1) - \psi(H_2)] + c_o''\psi(O_{2py})$$

The next step is to use the molecular orbitals from this symmetry analysis as the basis for a self-consistent field calculation on the water molecule. This will yield a set of coefficients for the molecular orbitals and a molecular orbital energy level diagram. These results are given below. You should compare these results with the traditional valence bond (Lewis structure) formulation of the bonding in H₂O.

$$\Psi(A_1) = 0.15[\psi(H_1) + \psi(H_2)] + 0.82\psi(O_{2s}) + 0.13\psi(O_{2p_z})$$

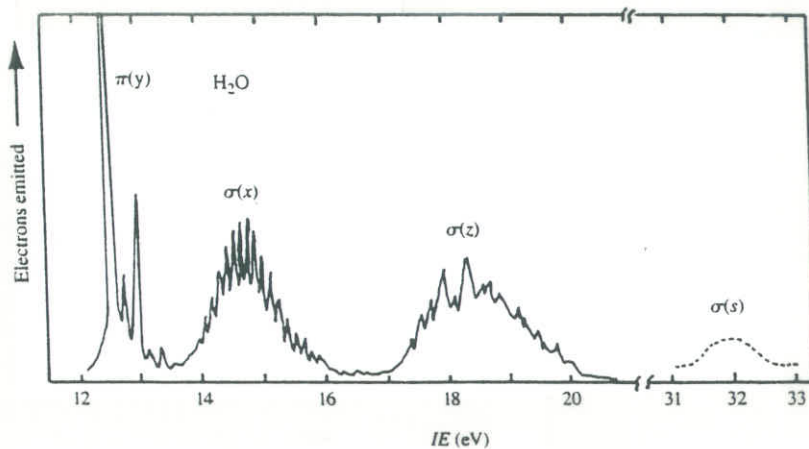
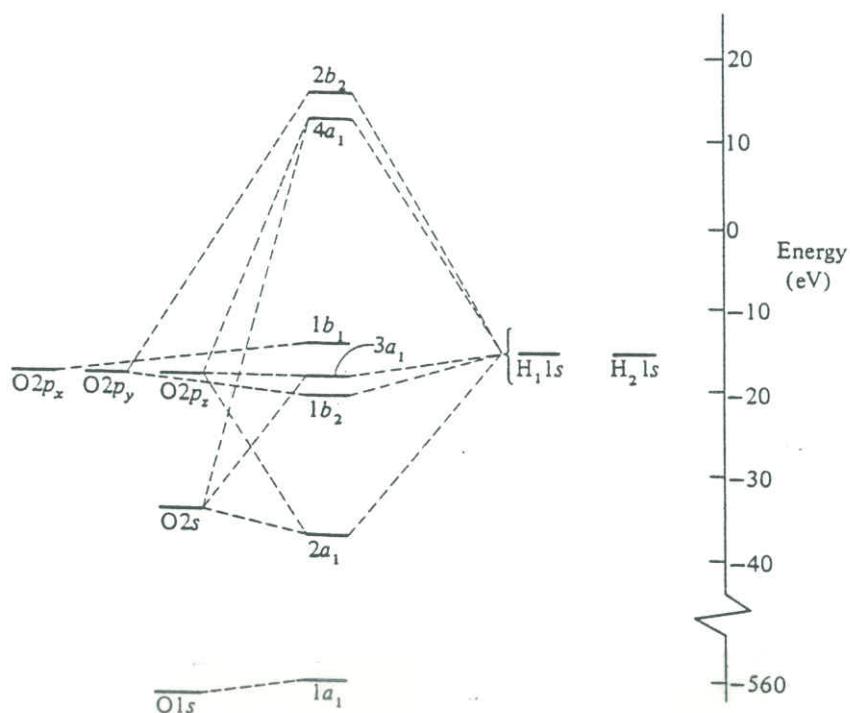
$$\Psi(B_2) = 0.42[\psi(H_1) - \psi(H_2)] + 0.62\psi(O_{2p_y})$$

$$\Psi(A_1) = 0.26[\psi(H_1) + \psi(H_2)] - 0.50\psi(O_{2s}) + 0.79\psi(O_{2p_z})$$

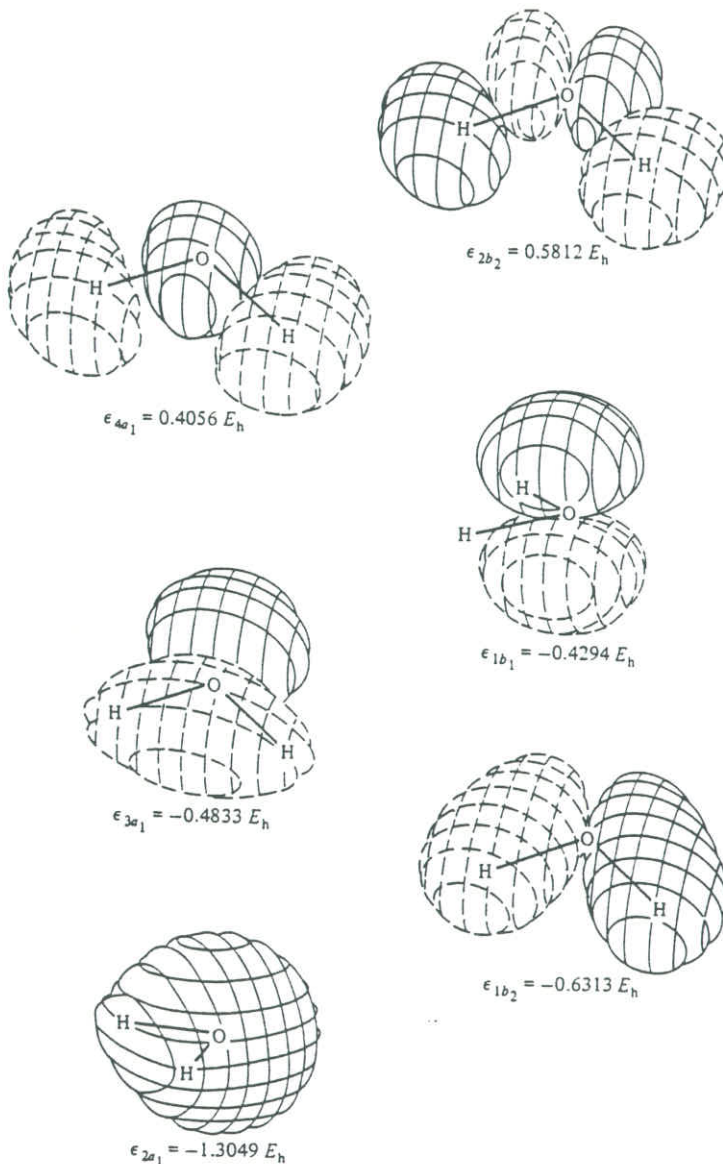
$$\Psi(B_1) = \Psi(O_{2p_x})$$

$$\Psi(A_1) = 0.75[\psi(H_1) + \psi(H_2)] - 0.84\psi(O_{2s}) - 0.70\psi(O_{2p_z})$$

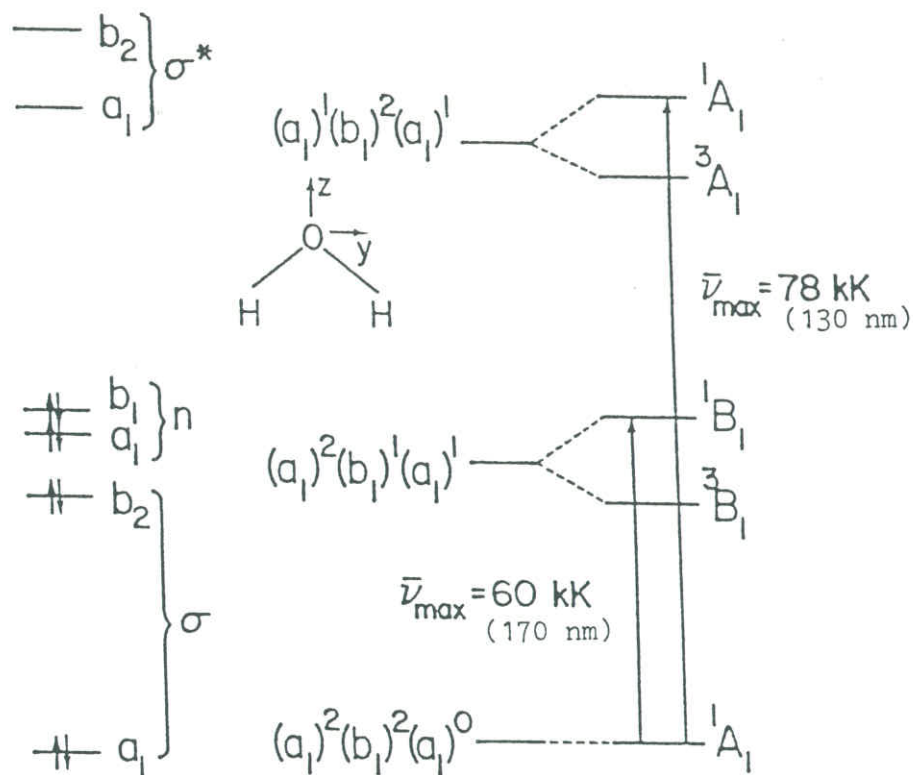
$$\Psi(B_2) = 0.89[\psi(H_1) - \psi(H_2)] - 0.99\psi(O_{2p_y})$$



The $1a_1$ MO is, as would be expected, the oxygen non-valence $1s$ orbital. The $2a_1$ MO is mainly a linear combination of the hydrogen $1s$ orbitals and the $2s$ on oxygen. The $1b_2$ MO is entirely the oxygen $2p_y$ and the hydrogen $1s$ atomic orbitals. The $3a_1$ MO and the $1b_1$ MO correspond to the two lone pairs. The $3a_1$ MO consists mainly of the oxygen $2s$ and $2p_z$ and the hydrogen $1s$ atomic orbitals and has its maximum value in the $+z$ direction. The $1b_1$ is entirely oxygen $2p_x$ and thus is directed at right angles to the plane of the molecule; the lone pairs thus have their maxima in the xz plane.



The first two electronic transitions in water occur at 170 and 130 nm. The molecular orbital diagram obtained previously and the principles of group theory can be used to interpret the electronic spectrum as is shown in the figure below.



Molecular orbital scheme and state diagram for H₂O.

By analogy with IR spectroscopy those electronic transitions are allowed for which the transition moment integral is non-zero.

$$M = e \int \Psi^f(\text{elec}) \hat{q} \Psi^i(\text{elec}) d\tau$$

As in the case of IR spectroscopy, this integral can be non-zero if and only if $\Gamma_{\text{elec}}^f \cdot \Gamma_q \cdot \Gamma_{\text{elec}}^i$ is equal to or contains the totally symmetric irreducible representation, A_1 . The figure shows that the electronic ground state has A_1 symmetry while the electric dipole operator ($\mu = e q$) has the symmetry of the cartesian coordinates (A_1 , B_1 , or B_2). Thus $\Gamma_{\text{elec}}^f \cdot \Gamma_q \cdot \Gamma_{\text{elec}}^i$ will equal A_1 if the excited state has A_1 , B_1 , or B_2 symmetry. If this is the case then the transition is said to be orbitally allowed. Note that several orbitally allowed transitions are spin-forbidden (involve a change in spin multiplicity) and therefore do not occur.